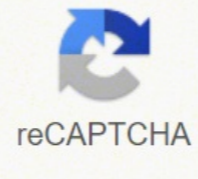




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## Which has permanent dipole moment

Something went wrong. Wait a moment and try again. Even though the total charge on a molecule is zero, the nature of chemical bonds is such that the positive and negative charges do not completely overlap in most molecules. Such molecules are said to be polar because they possess a permanent dipole moment. A good example is the dipole moment of the water molecule. Molecules with mirror symmetry like oxygen, nitrogen, carbon dioxide, and carbon tetrachloride have no permanent dipole moments. Even if there is no permanent dipole moment, it is possible to induce a dipole moment by the application of an external electric field. This is called polarization and the magnitude of the dipole moment induced is a measure of the polarizability of the molecular species. Dipole field Dipole potential In contrast to the permanent dipole of hydrogen bonds, the van der Waals bonds amount to a fluctuating dipole generated by the mutual influence of the motion of electrons of atoms. From: Cohesive Sediments in Open Channels, 2009

Acetylcholinesterase (AChE) from krait (*Bungarus fasciatus*) venom is a soluble, nonamphiphilic monomer of 72 kDa. This snake venom AChE has been analyzed by measurements of the stationary and the transient electric dichroism at different field strengths. The stationary values of the dichroism are consistent with the orientation function for permanent dipoles and are not consistent with the orientation function for induced dipoles. The permanent dipole moment obtained by least-squares fits for a buffer containing 5 mM MES is 1000 D, after correction for the internal directing field, assuming a spherical shape of the protein. The dipole moment decreases with increasing buffer concentration to 890 D at 10 mM MES and 770 D at 20 mM MES. The dichroism decay time constant is 90 ns ( $\pm$  10%) which is clearly larger than the value expected from the size/shape of the protein and indicates contributions from sugar residues attached to the protein. The dichroism rise times observed at low field strengths are larger than the decay times and, thus, support the assignment of a permanent dipole moment, although it has not been possible to approach the limit where the energy of the dipole in the electric field is sufficiently low compared to  $kT$ . The experimental value of the permanent dipole moment is similar to that calculated for a model structure of *Bungarus fasciatus* AChE, which has been constructed from its amino and acid sequence, in analogy to the crystal structure of AChE from *Torpedo californica*. Moderators: Chem\_Mod, Chem\_Admin 305174946 Posts: 61 Joined: Fri Sep 28, 2018 12:17 am Postby 305174946 » Thu Nov 08, 2018 5:21 pm Can someone explain permanent dipole moments to me please? As well as their importance 105114680 Posts: 60 Joined: Fri Sep 28, 2018 12:23 am Postby 105114680 » Thu Nov 08, 2018 5:41 pm Permanent Dipole Moments occur when the difference in the electronegativity of two atoms in a molecule is large. The more electronegative atom will attract more electrons, thus creating a partial negative charge around that atom and a partial positive charge around the atom with the lower electronegativity. Molecules with permanent dipole moments are just polar molecules, as opposed to having dipole moments that are induced by other molecules. They are important because the polarity of a molecule determines many of its properties and how it interacts with other molecules. Hope this helps! Jessica Chen 1F Posts: 61 Joined: Fri Sep 28, 2018 12:17 am Postby Jessica Chen 1F » Thu Nov 08, 2018 5:43 pm A permanent dipole moment is when the molecule is polar. The two or more atoms within the molecule must have substantially different electronegativities (one must attract electrons more than the other and becomes more negative while the other becomes positive). 005115864 Posts: 64 Joined: Fri Sep 28, 2018 12:15 am Postby 005115864 » Thu Nov 08, 2018 6:21 pm A permanent Dipole moment is when the difference of electronegativity between two atoms is so large that it holds a partial positive and partial negative charge. Hope this helps! :) Return to "Dipole Moments" Jump to Users browsing this forum: No registered users and 0 guests Tardigrade - CET NEET JEE Exam App © 2022 Tardigrade®. All rights reserved

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